

1. What are the 4 objectives of this section?
2. When do atoms bond?
3. Why do atoms bond?
4. What does the atoms electronic structure look like after they bond?
5. What pulls hydrogen atoms together?
6. How is a hydrogen molecule like helium?
7. What is wrong with using sticks to represent bonds?
8. How do ionic bonds form?
9. What types of atoms form cations and anions?
10. What does it mean to transfer electrons?
11. Why do they transfer electrons?
12. Why do they stick together after transferring electrons?
13. Why don't we talk about a molecule of salt?
14. What does salt's chemical formula tell us?
15. Why don't solid ionic compounds conduct electricity?
16. Why do ionic compounds conduct when dissolved?

17. Describe three properties of metals.
18. Why can metals conduct electricity?
19. Why are metals flexible?
20. What is a covalent bond?
21. What types of atoms form covalent bonds?
22. Why don't molecules conduct electricity when dissolved?
23. Why don't chlorine atoms transfer electrons in forming Cl_2 ?
24. Why do chlorine atoms share electrons?
25. What do you draw to indicate that the atoms are sharing electrons?
26. How do you form a double bond?
27. How do you form a triple bond?
28. How are double and triple bonds different from each other?
29. How do you form a polar covalent bond?
30. How do you know if the atoms are going to be attracted?
31. What is a polyatomic ion?
32. Why do you use parentheses around poly atomic ions?
33. What does the charge on the polyatomic mean?
34. What changes between the *-ate* and *-ite* polyatomic ions?