

Answer Questions 1-10 on page 101

1. _____

3. _____

5. _____

7. _____

9. _____

2. _____

4. _____

6. _____

8. _____

10. _____

Vocabulary

_____ 1. mole

_____ 2. Avogadro's constant

_____ 3. Molar mass

_____ 4. Conversion Factor

_____ 5. Metals

_____ 6. Nonmetals

_____ 7. Alkali metals

_____ 8. Alkaline earth metals

_____ 9. Transition metals

_____ 10. Halogens

_____ 11. Noble gases

_____ 12. Periodic law

_____ 13. Period

_____ 14. Group

_____ 15. Ionization

_____ 16. Ion

_____ 17. Cation

_____ 18. Anion

_____ 19. Atomic number

_____ 20. Mass number

_____ 21. Isotopes

_____ 22. Atomic mass unit

_____ 23. Average atomic mass

_____ 24. Nucleus

_____ 25. Proton

_____ 26. Neutron

_____ 27. Electron

_____ 28. Energy level

_____ 29. Orbital

_____ 30. Valence electron

A. The unit for the amount of a substance

B. The number of particles in 1 mole

C. The mass of one mole in grams

D. A ratio that has the same amount written two different ways

E. Elements that are malleable and ductile

F. Elements that are brittle and have dull luster

G. The elements in Group 1

H. The elements in Group 2

I. The elements in Group 3-12

J. The elements in Group 17

K. The elements in Group 18

L. When in atomic number order the properties of the elements repeat.

M. Columns on the periodic table.

N. Rows on the periodic tables.

O. The process of gaining or losing electrons.

P. An atom or group of atoms with a charge.

Q. Positively charged ion

R. Negatively charged ion

S. The number of protons in an atom

T. The total of protons and neutrons in an atom

U. Atoms with the same atomic number but different atomic numbers

V. One twelfth the mass of a carbon 12 atom

W. The mass of the average of the isotopes of an element

X. The dense, positive center portion of the electrons

Y. Positively charged pieces of atoms

Z. Neutral pieces of atoms

AA. Negatively charged pieces of atoms

BB. The path that an electron takes around the nucleus

CC. A region where you have a chance of finding electrons

DD. Electrons in the outermost energy level of an atom

12. Complete the following table.

Element	Symbol	Number of Protons	Number of electrons	Number of neutrons	Atomic Number	Mass Number
		25				53
			11	12		
		35		45		
					39	89
			33			75
	Ac					227

13. What was Dalton's atomic theory?

14. Describe Bohr's model of the atom.

15. Why is the periodic table shaped the way it is?

16. How do you find metals, and nonmetals on the periodic table?

17. Robyn recycled 15.1 mol of aluminum last month, how many grams of aluminum is this?

18. James has a balloon with 0.54 g of helium. How many moles of helium is this?